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⑤④ Absorbent materials, process for making them and their use in immobilising blood.

⑤⑦ An absorbent material useful for immobilising blood consists of a blood-swellable, substantially water-insoluble, covalently cross-linked anionic polyelectrolyte having at least some of its anions associated with transition metal counterions which are capable of coagulating blood, the polyelectrolyte having the property that when the polyelectrolyte is contacted with blood transition metal counterions dissociate from the polyelectrolyte, diffuse into the blood and cause coagulation thereof.

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**TITLE MODIFIED**  
**see front page**

ABSORBENT MATERIALS, PROCESS FOR MAKING THEM  
AND THEIR USE FOR IMMOBILISING BLOOD

This invention relates to absorbent materials and more particularly to absorbent materials for inclusion in products intended for the absorption of blood (which is to be understood herein as including menstrual fluid), such as sanitary  
5 towels and tampons, and in various other disposable absorbent products including those for medical and surgical usage. The invention also relates to processes for making the absorbent materials and to their use for immobilising blood.

In recent years a number of synthetic and partly  
10 synthetic water-swellaable, substantially water-insoluble, cross-linked polyelectrolytes have been proposed for use in disposable products because of their ability to absorb substantial amounts of body fluids, particularly water, urine

and blood. These absorbent materials are able to bind body fluids irreversibly so that absorbed fluid is rendered immobile and cannot be squeezed out under the pressures normally associated with the use of absorbent disposable products. While many such highly absorbent products can be used satisfactorily for the immobilisation of water, they are much less effective for immobilising urine and blood. In addition, while the absorbents can absorb the aqueous phase of blood and the lower molecular weight components thereof, blood also contains a high concentration of high molecular weight proteins which are not capable of being absorbed by these absorbent materials. Consequently, in a product containing the absorbent material, some of the blood tends to remain at the outside of a mass of the absorbent particles and in a fluid or mobile state. This fluid blood material can therefore leak from an absorbent product or cause staining, both of which are disadvantageous.

It is an object of the invention to provide an absorbent material particularly suitable for immobilising blood.

According to the invention there is provided a blood swellable, substantially water-insoluble, covalently cross-linked anionic polyelectrolyte having at least some of its anions associated with transition metal counterions which are capable of coagulating blood, the polyelectrolyte being characterised by the property that when the polyelectrolyte is contacted with blood transition metal counterions dissociate from the polyelectrolyte, diffuse into the blood

and cause coagulation thereof. This coagulation, which we believe to be due to the cross-linking of blood proteins by the transition metal counterions, is to be distinguished from the well known clotting process which is the normal  
5 response to wounding and which is initiated by calcium ions.

The anionic polyelectrolyte is preferably an anionic polysaccharide, for example an anionic derivative of starch, dextran or cellulose. While the transition metal ions will effect ionic cross-linking of anionic groups, the  
10 polyelectrolyte is covalently cross-linked whereby even in the absence of the transition metal ions the anionic polyelectrolyte is at least partially water-insoluble, eg at least 40% by weight water-insoluble. In a preferred form, the unmodified absorbent has a water-insolubility of  
15 at least 90% by weight. The covalent cross-linking may be effected by internal esterification as in US Patent Specification No. 3 678 931 (Buckeye Cellulose Corporation) or by suitable cross-linking agents which are well known to those in the art and include those referred to in German  
20 Patent Application No. P 2 702 781.2 (Unilever), German Patent Application No. P 2 634 539.1 (Hoechst AG), British Patent Specification No. 936 039 (Pharmacia) and US Patent Specification No. 3 589 364 (Buckeye Cellulose Corporation). Such cross-linking agents include the bifunctional compounds  
25 Q-R-Y where R is an alkylene group containing 1 to 10, and especially 2 to 5, carbon atoms, which may be substituted by one or more hydroxy groups, and/or which may be interrupted by heteroatoms, eg oxygen atoms, and Q and Y each represent

a halogen atom or an epoxy oxygen atom, the latter being linked to carbon atoms of the group R with the formation of an oxirane ring. Examples of such agents are epichlorhydrin, dichlorohydrin, dibromohydrin, 1,2-3,4-  
5 diepoxybutane, 1,2-7,8-diepoxyoctane, bisepoxypropylether and 1,4-butane-diol-bis-epoxypropylether.

The anionic groups are preferably carboxyalkyl groups, especially carboxymethyl or carboxyethyl groups. Poly-saccharide compounds preferably have a degree of substitution  
10 of anionic groups of at least 0.1, more preferably at least 0.2.

In the absorbent material of this invention at least some of the anions of the anionic polyelectrolyte are associated, or ionically complexed, with a transition metal  
15 counterion having special properties. Transition elements include elements 21 to 30 (scandium to zinc), 39 to 48 (yttrium to cadmium) and 57 to 80 (lanthanum to mercury) and the transition metal counterion present in the absorbent material of this invention is one which has the property of  
20 being able to cause the coagulation of blood. Transition metal ions effective for this purpose are known in the art and include especially copper and zinc ions although others such as nickel ions have this property to some extent. The  
25 bringing together of these coagulating ions and blood when using the absorbent material of this invention occurs simultaneously with the removal of part of the aqueous absorbable fraction of blood by the absorbent. By this means the immobilising action of the transition metal ions

on the unabsorbed blood proteins by the coagulant effect is enhanced. Consequently, the absorbent material of the invention immobilises blood by two distinct mechanisms:

1) the absorption by the absorbent of an aqueous fraction  
5 containing low molecular weight components; and 2) the gelation of a high molecular weight protein fraction which is not absorbed. A further factor contributing to the effectiveness of the absorbent material of this invention in the immobilisation of unabsorbed blood is the binding  
10 action by the gel formed by the association of the blood proteins with the transition metal ions on the residual aqueous fluids.

It is also to be appreciated that it is an essential feature of the absorbent material of the present invention  
15 that the blood-coagulating transition metal ions are present in such manner that when the absorbent material is contacted by blood they can diffuse from the polymer and cause gelation of blood proteins. Therefore the anionic groups of the poly-electrolyte must not bind too tightly the transition metal  
20 ions and the polymer network should not prevent the diffusion of the ions from the absorbent material to effect the desired coagulation of blood. The coagulant effect of the transition metal ions diffusing into blood is readily observable in the microscope as is more particularly described hereinafter.

25 Absorbent materials in accordance with this invention will generally contain at least 0.5 milliequivalents (meq) of the blood coagulating metal per gram, for example 0.5 to 10 meq/g.

It is not essential that all the anions of the polyelectrolyte be associated with the transition metal ions having a blood coagulating effect. Some of the anions may be associated with other counterions, for example  
5 alkali metal, alkaline-earth metal, ammonium or substituted ammonium ions, for example sodium, potassium, calcium, ammonium and tetramethylammonium ions.

Preferred absorbent materials of this invention are the zinc and copper salts of water-swellaable, covalently cross-  
10 linked carboxymethyl starch, dextran or cellulose.

Of special usefulness are the zinc salt forms of the absorbent materials described in German Patent Application No. P 2,702,781.2. However, also of particular usefulness are the zinc salt forms of the cross-linked carboxymethyl  
15 starch-, dextran- and cellulose-based water-swellaable absorbent materials described in the Hoechst, Pharmacia and Buckeye Cellulose Corporation patent specifications referred to above.

The absorbent product of the invention is preferably in particulate form such as in the form of a powder, granules,  
20 flakes or fibres.

The invention also relates to processes for making the absorbent materials of this invention.

Accordingly the invention also relates to a process for making an absorbent material of the invention in which  
25 process a water-swellaable, at least partially water-insoluble, covalently cross-linked anionic polyelectrolyte is treated in a swollen state with a compound of a transition metal whose ions are capable of coagulating blood so that at least

some of the anions of the polyelectrolyte become associated with transition metal counterions, and the treated polyelectrolyte is then dried.

5 According to one such process a water-swella-  
ble, covalently cross-linked, at least partially water-insoluble,  
anionic polyelectrolyte in its acid form is contacted in  
a water-swollen state with a compound of said transition metal  
reactive with the acidic groups of said polyelectrolyte to  
form metal salt groups; any non-volatile anions of the  
10 metal compound are removed by washing, and the water-  
swollen absorbent then dried. In this process the  
transition metal compound may be water-soluble or water-  
insoluble. Examples of suitable compounds are the basic  
zinc and copper compounds such as the carbonates, oxides  
15 and hydroxides of these metals.

The acid form of the polyelectrolyte may be obtained by the acid washing of a salt form.

In an alternative process a water-swella-  
ble, covalently cross-linked, at least partially water-insoluble, anionic  
20 polyelectrolyte in the form of its alkali metal, alkaline-  
earth metal, ammonium or substituted ammonium salt is  
contacted with a medium in which the polyelectrolyte is  
swella-ble and in which a soluble salt of said transition  
metal is dissolved whereby ion-exchange occurs and ions of  
25 said transition metal are introduced into the swollen  
polyelectrolyte, whereafter the polyelectrolyte is separated  
from said medium, washed and then dried. The swelling



medium for the polyelectrolyte is preferably water but suitable aqueous alcoholic media if they contain sufficient water to permit the polyelectrolyte to swell can also be employed; examples of alcohols for such media are methanol, ethanol and propanol. Examples of suitable transition metal salts for use in this process are the chlorides, nitrates, sulphates and acetates of copper and zinc although any other substantially water-soluble salt of the transition metal can be used.

10       The introduction of the transition metal ions into the absorbent material in the above processes will, where they replace monovalent cations, act to further cross-link temporarily the polyelectrolyte and this will enhance the wet-out of a mass of the absorbent by blood. However, if  
15       desired, the absorbent may also be surface-treated to improve blood wet-out with a polyether as described in British Patent Application Serial No. 2 007 998 or with one of the hydrocarbon materials described in European Patent Publication No. 9977.

20       The invention also relates to the use in absorbent articles for immobilising blood of the absorbent material of the invention. The absorbent article may comprise a fibrous carrier or support for the absorbent material, such as a woven or unwoven material such as cotton cloth, rayon,  
25       wool, surgical gauze or paper as well as cellulosic fluff, on or within which the absorbent material is supported. The absorbent material may be spread on the carrier or it may be mixed with loose fibres to make a composite fluff

or wadding which can be enclosed between cover sheets of paper or cloth. The article may also be in the form of a laminate. In a particular form, the carrier comprises two sheets between which the absorbent material is  
5 sandwiched.

The Applicants are aware that it has already been proposed in US Patent Specification No. 4 043 952 (National Starch and Chemical Corporation) to surface-treat a particulate water-swellable anionic polyelectrolyte in a  
10 medium in which the polyelectrolyte does not swell with a zinc salt to ionically complex the particle surface by zinc ions. The water-swellable, anionic polyelectrolytes surface-treated in said US Patent Specification No. 4 043 952 are of three kinds, viz water-soluble materials;  
15 water-insoluble, covalently cross-linked materials; and water-insoluble, ionically complexed materials. As described hereinafter, we have surface-treated various absorbents in accordance with the procedure described in said patent specification with a zinc compound but the  
20 treated absorbents have not displayed the critical property characteristic of the absorbent materials of this invention, that is that the blood-coagulating ion dissociates from the polyelectrolyte, diffuses into blood and causes coagulation thereof.

The invention will now be illustrated by the following Examples of which Examples 1 to 10 describe the preparation of absorbents according to the invention whereas Examples A to E concern absorbents not according to this invention but which are included for comparative purposes. Percentages are by weight unless otherwise specified.

Example 1

Potato starch (1000g) was slurried in water (950 ml) containing epichlorohydrin (8.4 ml; 1.0% epichlorohydrin by weight of starch). Sodium hydroxide (5g) in water (50 ml) was added with stirring and the mixture was applied to a heated roller via a feeder roller to form a layer on the surface of the roller of about 0.5 mm thickness. The roller itself was heated using steam at 40 lbs/sq. in gauge (3.77 bars) (140°C). The cross-linked starch derivative was removed from the roller as a flake material.

Sodium hydroxide (34g) in water (66 ml) followed by monochloroacetic acid (39 g) in water (11 ml) was slowly added with stirring to the cross-linked potato starch (100g) as prepared above. The mixture was aged overnight in a polythene bag. The theoretical degree of substitution was 0.67.

The moist carboxymethyl derivative was dispersed in 10 times its weight of 1N hydrochloric acid and soaked for 15

min. then filtered. The gel cake was repeatedly dispersed in water and filtered until the filtrate was substantially free of chloride ions. Ammonium hydroxide, specific gravity 0.910, (70 ml) was mixed with the water swollen washed acid  
5 cake before drying in a forced air oven (70°C).

The absorbent (50g) was soaked in an aqueous solution of zinc nitrate (2½ l, 1M) overnight, filtered, washed until the filtrate was substantially free of nitrate ions then dried in a forced air oven at 70°C and milled through a 2mm screen.

10

#### Example 2

Example 1 was repeated as far as ageing the carboxy-methylated mixture in a polythene bag.

The moist carboxymethyl derivative was repeatedly dispersed in water and filtered until the filtrate was neutral.  
15 The gel cake was then dispersed in an aqueous solution of copper sulphate (5 l, 1M), soaked overnight, then washed and filtered until the filtrate was substantially free of sulphate ions, and dried in a forced air oven at 70°C.

#### Example 3

20 Example 1 was repeated except that after washing of the acid cake, basic zinc carbonate (10.8g, 100% neutralisation) was thoroughly stirred into the acid cake and the product dried in a forced air oven at 70°C.

#### Example 4

25 Example 3 was repeated but zinc oxide (7.75g, 100% neutralisation) was stirred into the acid cake before drying.

Example 5

Example 3 was repeated except enough basic zinc carbonate was added to the cake to give theoretically 80% neutralisation of the acid groups, the remainder being  
5 neutralised by stirring in the appropriate quantity of ammonia solution (specific gravity 0.910).

Example 6

Example 1 was repeated but the absorbent (100g) was soaked in nickel chloride solution (5 l, 1M) overnight,  
10 filtered, washed until the filtrate was substantially free of chloride ions, then dried in a forced air oven at 70°C.

Examples 7-10

Absorbent (10g), as detailed below, was soaked in copper sulphate solution (500 ml, 1M) overnight, washed until  
15 the filtrate was substantially free of sulphate ions, dried in a forced air oven at 70°C and then milled through a 2mm screen. The following absorbents were used in these Examples:

	<u>Example</u>	<u>Absorbent</u>
20	7	P
	8	Q
	9	R
	10	S

Absorbent P was a carboxymethylated cellulose cross-linked by  
25 intermolecular esterification and available from Hercules Corporation under the trade name "Aqualon R" and generally described by Podlas, T.J. in INDA Tech. Symp. 1976, 2-3 March, pp 25-39.

Absorbent Q was a cross-linked sodium carboxymethyl cellulose obtained from Courtaulds Limited under the name of "Superabsorbent Blend DP 1006".

5 Absorbent R was a cross-linked sodium carboxymethyl cellulose obtained from Hoechst AG under the designation "2608/011".

Absorbent S was a cross-linked carboxymethyl dextran obtained from Pharmacia under the name of "CM-Sephadex C-50".

Comparative Examples A to C

10 Absorbent (20g), as detailed below, was added to a 500 cm<sup>3</sup> conical flask equipped with stirrer, water bath and condenser. Methanol (75 ml) was added followed by sufficient zinc acetate to give a cation level of 10 milliequivalents Zn<sup>2+</sup> per gram of absorbent on a dry weight basis. The  
15 mixture was stirred and heated to 68°C for 1 hour before cooling, filtering and drying.

Comparative Examples A and B illustrate the surface complexing of water-insoluble, covalently cross-linked materials and Comparative Example C illustrates the surface  
20 complexing of a water-insoluble, ionically complexed material.

The following absorbents were used in these Examples:

	<u>Comparative Example</u>	<u>Absorbent</u>
	A	T
	B	U
25	C	V

Absorbent T is the material as prepared in Example 1 but without the zinc treatment.

Absorbent U was a hydrolysed starch-polyacrylonitrile graft copolymer available from General Mills Inc., under  
5 the trade name "SGP-502S" and generally described in US Patent No. 3,997,484.

Absorbent V was the potassium salt of a polyacrylic acid cross-linked by aluminium ions available commercially from National Starch Corporation under the trade name  
10 "Permasorb 30" and generally described in US Patent No. 4,090,013.

#### Comparative Example D

Example 1 was repeated as far as ageing the carboxy-methylated mixture in a polythene bag. The moist carboxy-  
15 methyl derivative was repeatedly dispersed in water and filtered until the filtrate was neutral. The gel cake was then dried in a forced air oven at 70°C and milled through a 2mm screen.

#### Comparative Example E

20 Example 1 was repeated but instead of being treated with zinc nitrate the absorbent (100g) was soaked in aluminium sulphate solution (5l, 0.5M) overnight, washed and filtered until the filtrate was substantially free of sulphate ions, dried in a forced air oven at 70°C and milled through a 2mm  
25 screen.

All of the absorbents referred to in the above examples were, prior to treatment with the transition metal compound, insoluble in water to the extent of at least 50% by weight.

The treated materials were highly water-insoluble.

The blood coagulating property of the absorbents of Examples 1 to 10 and especially the contrast therewith of the effect of the absorbents of the Comparative Examples A to E was readily observed under the microscope. The examinations were made using a Leitz Ortholux II microscope under "dark field" or "phase contrast" conditions and a magnification of 250 times. The test fluid used was bovine serum albumin (5%) in sodium chloride solution (0.9%) rather than haemolysed blood because of its more constant composition, greater stability to microbiological attack and its optical transparency.

One or two drops of the bovine serum albumin test fluid were placed on a microscope slide and then a small quantity of the absorbent material placed on top, the whole being covered with a cover glass. The absorbent material was then studied for up to 1 hour.

Taking the results of all the examinations as a whole three types of effect could be clearly distinguished.

Effect 1 - There immediately begins to form an area of gelling protein adjacent to the particles, which area becomes both larger and denser with time.

Effect 2 - A slight precipitate was seen after about 30 minutes as a brightening of the surface of the particles, this probably being due to adsorbed protein.



Effect 3 - No precipitate observed.

Effect 1 is displayed by absorbents according to the invention.

Similar effects are obtained using serum prepared  
5 from blood.

The following Table 1 lists the effects produced by the materials of the various examples and indicates the nature of the counterion of the respective polyelectrolyte absorbent.

Table 1

10	<u>Example</u>	<u>Counterion</u>	<u>Type of Effect</u>
	1	$\text{Zn}^{2+}$	1
	2	$\text{Cu}^{2+}$	1
	3	$\text{Zn}^{2+}$	1
	4	$\text{Zn}^{2+}$	1
15	5	$\text{Zn}^{2+}/\text{NH}_4^+$	1
	6	$\text{Ni}^{2+}$	1
	7	$\text{Cu}^{2+}$	1
	8	$\text{Cu}^{2+}$	1
	9	$\text{Cu}^{2+}$	1
20	10	$\text{Cu}^{2+}$	1
	A	$\text{NH}_4^+ (\text{Zn}^{2+})^*$	2
	B	$\text{Na}^+ (\text{Zn}^{2+})^*$	2
	C	$\text{K}^+/\text{Al}^{3+} (\text{Zn}^{2+})^*$	2
	D	$\text{Na}^+$	3
25	E	$\text{Al}^{3+}$	3

$(\text{Zn}^{2+})$  indicates zinc ion at particle surface

Urine retention values and blood retention values for the materials of the above examples were determined by the following methods.

Determination of Urine Retention Value

5       The sample to be tested (0.20 g) is weighed out into a pre-weighed sintered glass Gooch crucible. Synthetic urine (5 ml) is added to the sample ensuring that the sample is completely wetted and it is left to soak for the prescribed time before being placed in a centrifuge tube and spun for 10 minutes at 850 rpm in a centrifuge having a head radius of 9 cm. The crucible with contents is then reweighed. The urine retention value is expressed as the weight of urine retained per gram of dry absorbent. The formula of synthetic urine, derived from the information given in the Handbook of Clinical Laboratory Data, 2nd Edition, 1968, 15 pages 17-20, is a solution in 5 litres of water of the following:

		<u>grams</u>
20	CaCl <sub>2</sub> ·2H <sub>2</sub> O	3.680
	K <sub>2</sub> SO <sub>4</sub>	0.175
	KCl	44.740
	KOH	2.190
	NH <sub>4</sub> Cl	6.020
	Citric acid	2.630

25       Experiments have shown that the same urine retention values are obtained by using natural urine in place of the synthetic urine.

Determination of Blood Retention Value

To determine the blood retention value, the sample to be tested (0.20g) is weighed out and sprinkled onto haemolysed blood (5-10 ml) in a 25 ml glass beaker. After leaving to  
5 soak for the prescribed time, the contents are transferred to a pre-weighed sintered glass Gooch crucible and placed in the centrifuge tube and spun in the same manner as described above. The crucible with contents is then reweighed. The blood retention value is expressed as the weight of blood  
10 retained per gram of dry absorbent.

Usually the blood and urine retention values of absorbent materials are very similar or equivalent but in the case of the absorbent materials of the invention the blood retention value rises with time and reaches a maximum within about 2  
15 hours. The maximum is occasionally reached in a considerably shorter time and hence the blood retention value measured after 10 minutes can be similar to that measured after 2 hours. Similar results are obtained using menstrual fluid in place of blood.

20 These retention measurements effectively measure the amount of material (urine or blood) rendered immobile under the conditions of the test. Consequently, the blood retention value will take account of blood which while not absorbed within the particles of the polyelectrolytic  
25 absorbent is coagulated and rendered immobile even under those conditions of the test which are designed to remove

absorbate between the particles of the absorbent. The higher blood retention values after the contact period of 2 hours for the absorbent materials of this invention are indicative of the increase in the degree of immobilisation of blood obtained with these materials.

During the determination of the blood retention value the swelling of the absorbent material may tend to block the pores in the sinter glass crucible and hinder non-bound fluid from draining during centrifugation. For this reason, the qualitative microscopic examination is a better way of determining or demonstrating a coagulating effect on blood of an absorbent containing metal ions. The microscopic examination method must also be relied upon to differentiate absorbents of the invention from those absorbents, such as Absorbent V, which are ionically cross-linked by polyvalent ions who do not have the property of coagulating blood, since these absorbents also have a swelling capacity on exposure to blood which increases with time (and therefore the blood retention value increases in time) due to diffusion of cross-linking ions from the absorbent particles.

The determined urine retention value (URV) and blood retention values (BRV) for the products of the above examples are given in Table 2 which also includes the number of milliequivalents of cation per gram of dry polymer. The retention values are given to the nearest whole number.

TABLE 2

	<u>Example</u>	<u>URV</u> (2hr soak)	<u>BRV</u> (10min soak)	<u>BRV</u> (2hr soak)	<u>meq</u> <u>Cation</u>	<u>Cation</u>
5	1	7	13	22	2.21	Zn
	2	2	21	32	2.79	Cu
	3	7	11	22	1.77	Zn
	4	6	13	31	2.11	Zn
10	5	7	13	36	2.09 0.33	Zn NH <sub>4</sub>
	6	5	7	11	2.26	Ni
	7	3	10	16	2.37	Cu
	8	1	6	9	4.86	Cu
	9	2	*	23	3.02	Cu
	10	2	4	29	3.49	Cu
15	A	9	10	10	0.26 0.16	NH <sub>4</sub> Zn <sup>4</sup>
	B	*	*	24	2.69 1.95	Na Zn
	C	14	15	29	5.41 0.17 0.25	K Zn Al
20	D	12	12	12	2.29	Na
	E	2	1	4	0.32	Al

\* value unobtainable due to poor drainage of the swollen material in the centrifugation.

25

## CLAIMS:

1. An absorbent material suitable for immobilising blood consisting of a blood-swella-  
ble, substantially water-insoluble, covalently cross-linked anionic  
5 polyelectrolyte having at least some of its anions associated with transition metal counterions which are capable of coagulating blood, the polyelectrolyte being characterised by the property that when the polyelectrolyte is contacted with blood transition metal counterions  
10 dissociate from the polyelectrolyte, diffuse into the blood and cause coagulation thereof.
2. An absorbent material as claimed in Claim 1, characterised in that the polyelectrolyte is an anionic  
15 polysaccharide.
3. An absorbent material as claimed in Claim 1 or Claim 2, characterised in that the anionic groups are carboxyalkyl groups.  
20
4. An absorbent material as claimed in any of Claims 1 to 3, characterised in that the transition metal is zinc or copper.
- 25 5. Process for making an absorbent material as claimed in Claim 1, characterised in that a water-swella-  
ble, at least partially water-insoluble, covalently cross-linked anionic polyelectrolyte is treated in a  
swollen state with a compound of a transition metal whose  
30 ions are capable of coagulating blood so that at least some of the anions of the polyelectrolyte become associated with transition metal counterions, and the treated polyelectrolyte is then dried.

6. A process as claimed in Claim 5, characterised in that the water-swella-  
ble, covalently cross-linked, at least partially water-insoluble, anionic polyelectrolyte in  
its acid form is contacted in a water-swollen state with  
5 a compound of said transition metal reactive with the  
acidic groups of said polyelectrolyte to form metal salt  
groups; any non-volatile anions of the metal compound  
are removed by washing, and the water-swollen absorbent  
then dried.
- 10 7. A process as claimed in Claim 6, characterised  
in that the transition metal compound is a basic  
zinc or copper compound.
- 15 8. A process as claimed in Claim 7, characterised  
in that the transition metal compound is zinc or copper  
carbonate, oxide or hydroxide.
- 20 9. A process as claimed in Claim 5, characterised  
in that the water-swella-  
ble, covalently cross-linked, at  
least partially water-insoluble, anionic polyelectrolyte  
in the form of its alkali metal, alkaline- earth metal,  
ammonium or substituted ammonium salt is contacted with  
a medium in which the polyelectrolyte is swella-  
25 ble and in  
which a soluble salt of said transition metal is dissolved  
whereby ion-exchange occurs and ions of said transition  
metal are introduced into the swollen polyelectrolyte,  
whereafter the polyelectrolyte is separated from said medium,  
washed and then dried.
- 30 10. A process as claimed in Claim 9, characterised in  
that the swelling medium for the polyelectrolyte is water.
- 35 11. A process as claimed in Claim 9 or 10, characterised in  
that the transition metal salt is zinc or copper chloride,  
nitrate, sulphate or acetate.

12. Use of an absorbent material as claimed in  
Claim 1 for immobilising blood.





European Patent  
Office

# EUROPEAN SEARCH REPORT

0019371

Application number

EP 80 30 1307

DOCUMENTS CONSIDERED TO BE RELEVANT			CLASSIFICATION OF THE APPLICATION (Int. Cl. 3)
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	
X	DE - A - 2 737 994 (UNION CARBIDE) * Page 4, claims 1 and 18; page 38, lines 3-7 * --	1,3,4,12	A 61 L 15/00
X	FR - A - 2 305 452 (SANYO) * Page 1, lines 27-34; page 8, lines 1-5; claims 1,2,8,9 * --	1-4,12	
X	FR - A - 1 293 415 (KIMBERLY-CLARK) * Page 12, lines 20-75; abstract * --	1,4,12	TECHNICAL FIELDS SEARCHED (Int.Cl. 3)
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<input checked="" type="checkbox"/> The present search report has been drawn up for all claims			
Place of search The Hague		Date of completion of the search. 28-08-1980	Examiner PELTRE

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